

















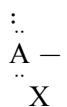










## Table Summary Animated - VSEPR

21:18:36 2025-02-02

A = Central Atom, X = an atom bonded to A, E = a lone pair on A

Total Domains	Geometric Formula	Picture	Bonded Atoms	Lone Pairs	Molecular Shape	Electronic Geometry	Example	Hybridization	Bond Angle
1	AX		1	0	Linear	Linear		$s$	$180^\circ$
2	AX <sub>2</sub>		2	0	Linear	Linear		$sp$	$180^\circ$
2	AXE		1	1	Linear	Linear		$sp$	$180^\circ$
3	AX <sub>3</sub>		3	0	Trigonal Planar	Trigonal Planar		$sp^2$	$120^\circ$
3	AX <sub>2</sub> E		2	1	Bent	Trigonal Planar		$sp^2$	$120^\circ$
3	AXE <sub>2</sub>		1	2	Linear	Trigonal Planar		$sp^2$	$120^\circ$
4	AX <sub>4</sub>		4	0	Tetrahedral	Tetrahedral		$sp^3$	$109.5^\circ$
4	AX <sub>3</sub> E		3	1	Trigonal Pyramidal	Tetrahedral		$sp^3$	$109.5^\circ$
4	AX <sub>2</sub> E <sub>2</sub>		2	2	Bent	Tetrahedral		$sp^3$	$109.5^\circ$
4	AXE <sub>3</sub>		1	3	Linear	Tetrahedral		$sp^3$	$109.5^\circ$

Total Domains	Geometric Formula	Picture	Bonded Atoms	Lone Pairs	Molecular Shape	Electronic Geometry	Example	Hybridization	Bond Angle
5	$AX_5$		5	0	Trigonal Bipyramidal	Trigonal Bipyramidal		$sp^3d$	$90^\circ$ , $120^\circ$
5	$AX_4E$		4	1	Seesaw	Trigonal Bipyramidal		$sp^3d$	$90^\circ$ , $120^\circ$
5	$AX_3E_2$		3	2	T-Shaped	Trigonal Bipyramidal		$sp^3d$	$90^\circ$ , $120^\circ$
5	$AX_2E_3$		2	3	Linear	Trigonal Bipyramidal		$sp^3d$	$90^\circ$ , $120^\circ$
6	$AX_6$		6	0	Octahedral	Octahedral		$sp^3d^2$	$90^\circ$
6	$AX_5E$		5	1	Square Pyramidal	Octahedral		$sp^3d^2$	$90^\circ$
6	$AX_4E_2$		4	2	Square Planar	Octahedral		$sp^3d^2$	$90^\circ$